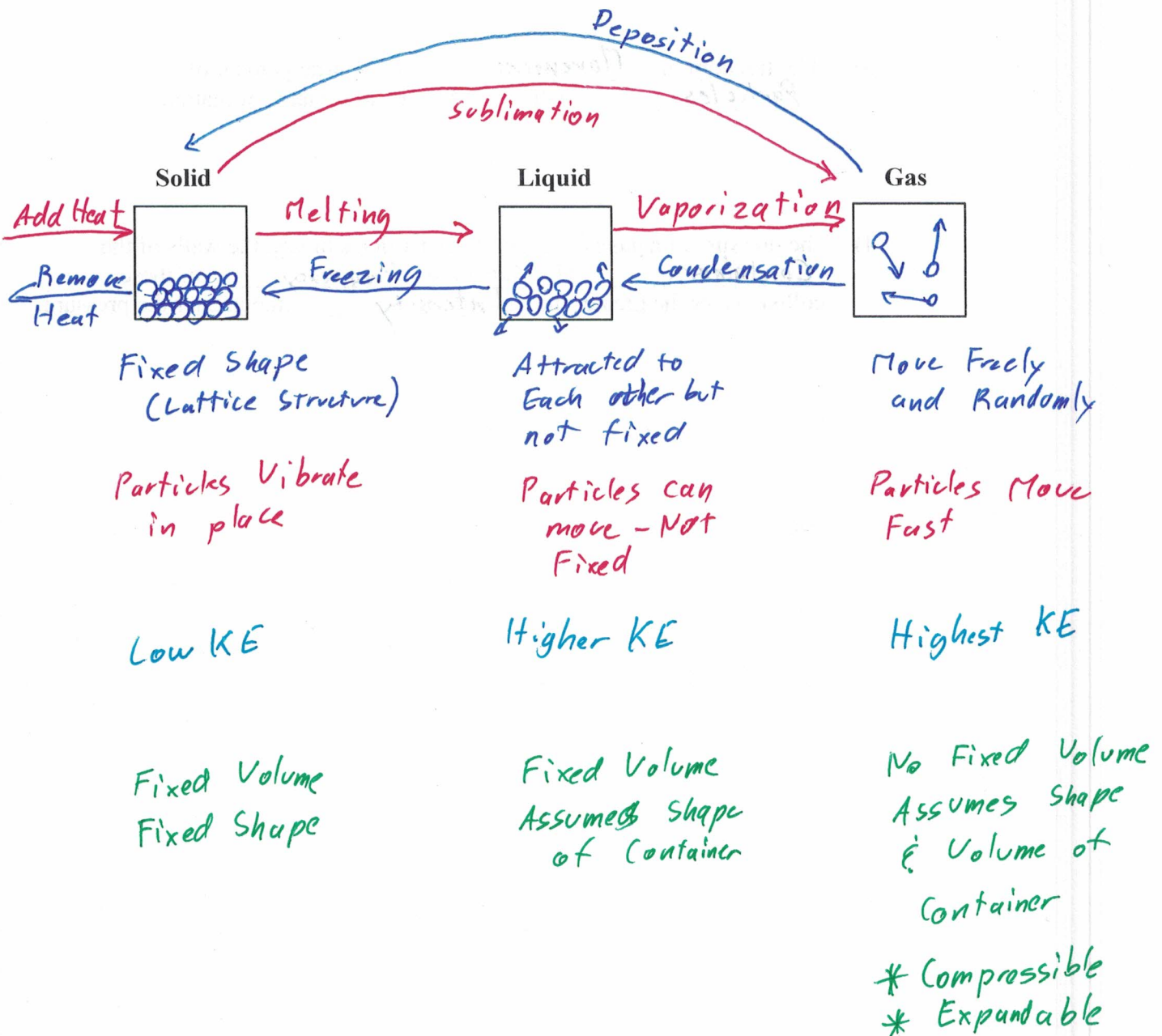


<b>The Kinetic Model of Matter</b>	Name:
	Period:
	Date: 8/20/18

**Kinetic Energy** – energy associated with Motion.

**Kinetic Energy** affects the State/Phase of a substance.

**Temperature** – the Average Kinetic Energy of a substance.

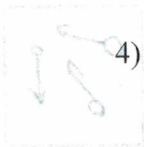


**KINETIC MODEL OF MATTER (Kinetic Theory)**

- 1) All matter is made up of Very small particles
- 2) The particles are always Moving.

Higher Temperature = Higher KE

- 3) The freedom of Movement and the arrangement of Particles is different for the three states of matter.



- 4) The pressure of a gas is a result of gas particles hitting the walls of the Container. The higher the Frequency of these collisions, or the greater the Intensity, the greater the pressure.

Handwritten notes in the bottom half of the page, including:

- Fixed shape (Surface structure)
- Particles vibrate in place
- Low KE
- Fixed Volume
- Assumes shape of the container
- Fixed Volume
- Higher KE
- Fixed Volume
- Highly KE
- Particle move fast with no attraction to each other
- Free Volume
- Free Volume
- \* Compressible
- \* Expandable
- no fixed volume
- Fill container